

Molarity & Titration

Calculate the molarity of a solution made by dissolving **15.0 g of KNO_3** in **200. mL of water**.

What is the molarity of a KOH solution if 28.0 mL neutralizes 15.0 mL of 0.195 M H_2SO_4 ?

What is the molarity of a $\text{Ca}(\text{OH})_2$ solution if 50.0 mL of $\text{Ca}(\text{OH})_2$ neutralizes 40.0 mL of 0.250 M HNO_3 ?

Calculate the molarity of the following solution and its ions: 15.6 g of potassium hydroxide (KOH) is dissolved in water to make 300 mL of solution.

What is the equation for dilution?

How many milliliters of 12.0 M HCl are needed to make 1.5 L of 0.50 M HCl?

200 mL of a 6.0 M HCl solution is diluted by adding water to make the total volume 1.0 L.
What is the new concentration of the HCl solution?

A 40.0 mL sample of a sulfuric acid (H_2SO_4) solution is titrated with 0.250 M sodium hydroxide (NaOH). It takes 50.0 mL of NaOH to reach the endpoint. What is the molarity of the sulfuric acid solution?